Electronic Supplementary Data

Kinetic study of oxidation of nitrite with a metallo superoxide

Sekhar Gain[a], Subrata Mukhopadhyay[b,*] & Rupendranath Banerjee[b,*]

[a]Ramakrishna Mission Vidyamandira, Department of Chemistry, Belur Math, Howrah 711 202, India

[b]Department of Chemistry, Jadavpur University, Kolkata 700 032, India

Email: smukhopadhyay@chemistry.jdvu.ac.in (SM)/ rupenju@yahoo.com (RB)

No. **Contents** **Pg. No.**

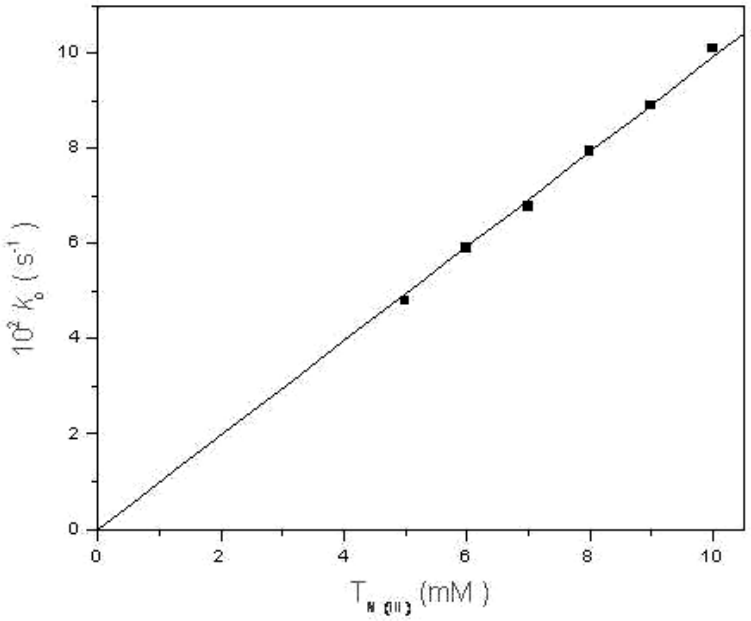

Fig. S1 – Variation of k_o with $T_{N(III)}$. [**(1)**] = 0.50 m***M***, ***p***H = 4.0, ***T*** = 25.0 °C.

Table S1 – Stoichiometric results for the oxidation of nitrite by **(1)** at ***p***H = 4.0

[**(1)**] (Mm)	$T_{N(III)}$ (mM)	Residual $T_{N(III)}$ (mM)	$\Delta[1]/\Delta T_{N(III)}$
1.0	5.0	4.5	2.0
2.0	10.0	9.1	2.2
4.0	15.0	12.9	1.9

Table S2 – Variation of k_o with nitrite. [**(1)**] = 0.5 m***M***, ***p***H = 4.0, ***T*** = 25.0 °C

$T_{N(III)}$ (mM)	$10^2 k_0$ (s^{-1})
5	4.8
6	5.9
7	6.8
8	7.9
9	8.9
10	10.1

Table S3 – Variation of k_o with ***p***H. [**(1)**] = 0.5 m***M***, $T_{N(III)}$ = 6m***M***, ***T*** = 25.0 ^oC

pH	10^2k_0 (s^{-1})
3.6	10.1
3.8	7.4
4.0	5.9
4.2	5.1
4.4	4.6
4.6	4.2
4.8	3.9

Table S4 **–** Effect of D_2O on k_o. [**(1)**] = 0.5 m***M***, $T_{N(III)}$ = 6m***M***, ***p***D = 4.4, ***T*** = 25 ^0C

% ***(v/v)*** of D_2O	10^2k_0 (s^{-1})
92.5	2.1
75	2.6
50	3.4
30	4.0
0	4.6